

## Section 71 Ionic And Metallic Bonding Answers

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## **7 1 Ion Answer Key**

The strength of metallic bonding is a function of the number of electrons provided by the atoms and the consequent charge on the metal ions. The ionic radius also plays a part, as smaller ions exert a greater force of attraction on the negative charge cloud. Increasing ionic charge = stronger metallic bonding

## **3.1.3 Bonding - Nature of ionic, covalent and metallic bonds**

Chapter 7 - Ionic and Metallic Bonding Section 7.3 - Bonding in Metals. The valence electrons of metal atoms can be modeled as a sea of electrons. Metallic bonds consist of the attraction of the free-floating valence electrons for the positively charged metal ions.

## **Chapter 7 Ionic And Metallic Bonding Practice Problems Answers**

7.1 Ion Formation 7.2 Ionic Bonds & Ionic Compounds 7.3 Names & Formulas for Ionic Compounds 7.4 Metallic Bonds & Properties of Metals

## **Chapter 7: Ionic Bonding (7.1) Flashcards | Quizlet**

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## 7 1 Ion Answer Key

Objectives<br />Describe the electron-sea model of metallic bonding, and explain why metals are good electrical conductors.<br />Explain why metal surfaces are shiny.<br />Explain why metals are malleable and ductile but ionic-crystalline compounds are not<br /> 3.

## Chapter 6.4 : Metallic Bonding - SlideShare

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## Chapter 7 Ionic And Metallic Bonding Practice Problems Answers

Section 7.4 Metallic Bonds and the Properties of Metals pages 225-228 Section 7.4 Assessment pages 228 40. Contrast the structures of ionic compounds and metals. Ions in ionic compounds are arranged in a repeating pattern of alternating

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charges, whereas metals consist of fixed cations surrounded by a sea of mobile, or delocalized, electrons. 41.

### **Ionic Compounds and Metals**

The lanthanide (/ ' l æ n θ ə n aɪ d /) or lanthanoid (/ ' l æ n θ ə n ɔɪ d /) series of chemical elements comprises the 15 metallic chemical elements with atomic numbers 57-71, from lanthanum through lutetium. These elements, along with the chemically similar elements scandium and yttrium, are often collectively known as the rare earth elements.. The informal chemical symbol Ln ...

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