

Chapter 11 Review Gases Answer Key

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Chapter 11 Review Gases Answer Key
CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression V = (a) increasing P (b) decreasing T

Home - Kenilworth Public Schools
If a gas and a liquid are the same temperature and pressure, diffusion occurs much faster in the gas because A. there are more elastic collisions between the particles in a gas B. gases are more compressible C. the particles move faster in a gas and there is a greater distance between them D. gas molecules are in continuous motion

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Ch. 11 Gas Laws Chapter Review Answers Particles of matter are Chapter 11 Test Review. multiple choice (25) definition & applications of pressure (also atmospheric) SI unit of force. definition & use of a barometer. standard temperature & pressure (STP) Definition of Dalton's law of partial pressures.

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CHAPTER . 11 . REVIEW . Gases . SHORT ANSWER . Answer the following questions in the space provided. 1. Consider the following data table: Approximate pressure (kPa) Altitude above sea level (km) 100 . o (sea level) 50 5.5 (peak of Mt. Kilimanjaro) 25 11 Get cruising altitude) < 0.1 22 (ozone layer) -. Kindle File Format Chapter11 Review Gases Answer Key

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Modern Chemistry 97 Gases CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. ____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. ____ For the expression , P

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CHAPTER 11 REVIEW . Gases . SHORT ANSWER . Answer the followin9 questions in the space provided. 1. c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of l mol. (c) multiplied by 22.4 . L. (b) divided by the mass of 1 mol. (d) divided by 22.4 . L. 2. c For the expression V = n–T.

CHAPTER REVIEW Gases
CHAPTER . 11 . REVIEW . Gases . SHORT ANSWER . Answer the following questions in the space provided. 1. Consider the following data table: Approximate pressure (kPa) Altitude above sea level (km) 100 . o (sea level) 50 5.5 (peak of Mt. Kilimanjaro) 25 11 Get cruising altitude) < 0.1 22 (ozone layer) -. .

REVIEW Gases
Chapter 11 Review Gases Section 1 Answers Modern Chemistry 97 Gases CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. ____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided ...

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Chapter 11 - Gases. Review Skills 11.1 Gases and Their Properties Ideal Gases Properties of Gases Discovering the Relationships Between Properties The Relationship Between Volume and Pressure. Internet: Boyle's Law.

Chapter 11 - Gases
CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: increase a. temperature increases, volume stays the same decrease b. volume increases, temperature stays the same

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CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Pressure surf f a o c r e ce area. For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged. 2. Rank the following pressures in increasing order. (a) 50 kPa (c) 76 torr (b) 2 atm (d) 100 N/m² 3.

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Chapter 11. Engine Fundamentals. 2. z. Identify the major parts of a typical automotive ... Cite and demonstrate safe working practices related to engines. z. Correctly answer ASE certification test questions that require knowledge of the basic operation of ... gases from the engine and readies the cylinder for a fresh charge of air and fuel.

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states that the rates of effusion of gases at the same temperature and pressure are inversely proportional to the square roots of their molar masses. Graham's law of effusion the process whereby the molecules of a gas are confined in a container randomly pass through a tiny opening in the container

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CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: increase a. temperature increases, volume stays the same decrease b. volume increases, temperature stays the same